# Pocl3 use in organic chemistry

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# Pocl3 use in organic chemistry

What is isopropyl in organic chemistry. What is the name of pocl3.

1.a. V. Simonyan, Khim.-factory. ZH., 33 (3), 43 - 44 (1999). Google ScholarÂ 2.a. A. Avanesyan and A. V. Simonyan, Khim.-factory. ZH., 35 (2), 37 - 38 (2001). Google ScholarÂ 3.M. D. Mashkovskii, Drugs, [Russian], Meditsina, Moscow (1986), Vols. 1 and 2. Google ScholarÂ 4.M. Boucard, H. Orzalesi, J. Refel and Y. Vedel, C. r. Acad. SCI., D267 (2), 244 - 246 (1968). Google ScholarÂ 5.B. Voigt, G. Wagner, P. Walsmann, et al., Pharmazie, 38 (12), 835 â € "838 (1983). PubMedÂ CasÂ Google ScholarÂ 6.us patent n. 3466308 (1967); Ref. Zh. Khim., 19n 166p (1970) .7.USC patent n. 4388459 (1980); Ref. Zh. Khim., 1O 52P (1984) .8.jpn. Patent No. 60-185752 (1984); Ref. Zh. Khim., 5N 115P (1987) .9.S. I. Lur'e and E. S. Chaman, Oxazolones, in: reactions and methods for investigation of organic compounds [in Russian], V. M. RODIONOVA, et al. (EDS), Goskhimizdat, Moscow (1959), Edition Izobret, No. 2 (1983). Paging 2 quote counts are provided by the web of science and the Crossref. Counts may vary by service and depend on the availability of their data. Counts will be updated daily once available. RITUSH KUMAR * AB AND_ RADHEY M. SINGHB Â Auction Affiliations An efficient Azaeterocicli etylated hydrochloride was obtained using POCL3 as a chlorination agent under metal-free reaction conditions. Mechanical studies show that the reaction proceeded through the nucleophile attack of Pocl3 on the nitrogen of the quinoline ring in a stereoselective way. The resulting products are intermediate versatile in organic synthesis and have been used in reaction synthesis and free of cross-coupling metals of heterocyclics. The developed protocol offers inexpensive materials and easily synthesized starting, easy operation and high efficiency with a high yield of products. This article is part of the themed collection: synthetic methodology in OBC You have access to this article Please wait for us to upload to your website ... Something went wrong. Try again? Additional Information PDF (10708K) Crystal CIF structure data (22K) Posted May 29, 2019 Org. Biomol. Chem., 2019,17, 5990-5996 R. Kumar and R. M. Singh, Org. Biomol. Chem, 2019, Â 17, 5990 DOI:. 10,1039 / C9OB00841A To request permission to reproduce material from this article, please go to the Copyright Clearance Center request page. If you are an author who contributes to an RSC publication, you do not need to request authorization provided that a correct recognition is provided. If you are the author of this article, you do not need to request permission to reproduce the figures and diagrams provided that a correct recognition is provided. If you want to reproduce the entire article in a third-party publication (excluding your thesis/disssertation for which no authorization is required), gorequest page of the copyright clearance center. Learn more about how to properly recognize RSC content. Tweet Share recovery data from CrossRef. This could take some time toLoading of related content Melting point: 1.25 °C (lit.) boiling point: 107 °C Density 1.645 g/mL to 25 °C (lit.) steam density 5.3 (vs air) steam pressure 104 mm Hg (50 °C) refractive index n20/D 1.461 Inflamability point: 105.8°C storage temperature. Preserve under +30°C. solubility Phosphorus(V) oxychloride is soluble in many organic solvents. shape Liquid Specific Gravity 1.692 (15/15°C) Color Color Color PH 1.0 (5g/l, H2O, 25°C) Odor Pungent odour Water Solubility reacts exoticly Sensitive Moisture Merck Sensitive 14,7349 Display Limits TLV-TWA0.628 mg/m3 (0.1 ppm) (ACGIH). Stability: stable. It reacts violently with water. Incompatible with many metals, alcohols, amines, phenol, DMSO, strong bases. PHYSTEM 1KHYPHYPHYPHYPHYPHYPHYPHYPHYPHYPHYPHYCYPHY It is a colorless and transparent liquid, and has an irritating unpleasant smell. It will smoke intensely in wet air. Its relative density is 1.68, melting point is 1.25°C, boiling point is 105.1°C. It breaks into phosphorous acid and hydrogen chloride in water and ethanol. When suddenly combined with a large amount of water, an intense reaction can occur. POCl3 reacts with water and alcohol to create phosphorous or phosphate acid. If alcohol replaces water in reaction, the final product will be the trialkyl phosphate. This type of reaction often occurs in pyridine or ammonia, since it absorbs the HCl produced to stimulate reaction. When catalyzed by a Lewis acid such as manganese chloride, POCl3 and a large amount of phenolic (ArOH) heats up to produce triaryl phosphate, as in the reaction below: 3 C6H5OH + O=PCl3 → O=P(OC6H5)3 + 3 HCl. Phosphorus oxychloride is a Lewis base, and compounds Cl3P5+O− + TiCl4 → Cl3P5+O-−TiCl4. Its adeduct with aluminum chloride (POCl3•AlCl3) is very stable, and therefore POCl3 is used to remove the AlCl3 in the final products of reactionsIn the presence of AlCl3, POCl3 reacts with hydrogen bromide to create POBr3. Phosphorus chloride can be used as a semiconductor dopant, andIt is a raw material for light conducting fibers. It is widely used in pesticides, pharmaceuticals, dyes, phosphates and flame retardant production. It is a raw material for the production of organic phosphorus herbicides and Chlordimeform, and is a plasticizer in the production of plastic. Phosphorus oxychloride is also used in the chlorination of long agitation sulfope drugs, is an intermediate in the production of dye and a catalyst in organic synthesis of chlorination agents, and is a uranium mine extraction agent. It is also used in the manufacture of pharmaceutical products. The toxicity is similar to phosphorus trichloride, phosphorus pentachloride and phosgene. Large mice, oral, LD50: 380 mg/kg; Inhaled, LC50: 32 PPM / 4H. Acute poisoning in small mice Result of restlessness, upper respiratory tract and conjunctival irritation, depression, convulsions, unstable walking, lying on the side, and eventually death. For large mice, in addition to the above symptoms, also exposed lacrimation, corneal fogging and pulmonary edema. Subacute and chronic toxicity: large mice, inhaled for 60 days at a concentration of 33.5 mg/m3, exhibited slowed weight gain, skin ulcers, reduced survival rates in macrophages of the lung, no liver and renal function and characteristic organ changes o. This product has a strong oxidizing and fat soluble effect; In addition to burning the digestive tract, it can also cause acute necrosis and autolysis in the liver when absorbed through the digestive tracts. In addition, in addition to burning the skin, when fully absorbed through the skin, 3% can cause death in animals. When phosphorus oxychloride is inhaled and comes into contact with the mucosa of the moist respiratory tract, it will break down into phosphoric acid and hydrogen chloride, and irritate and corrupt the mucosa. When humans contact each other at a concentration of 70 mg/m3, they usually show symptoms after a latency period of 2-6h, including irritation of the mucosa of the respiratory tract and eye pain. Severe cases include choking, cyanosis, pulmonary oedema, heart failure, anaemia, liver damage and proteinuria. When the temperature in the car is high and the humidity is relatively low, there is a high risk of inhalation and poisoning. The highest permissible concentration is 0.05 mg/m3. Operators should wear protective devices and production equipment should be tightly sealed. Processors can also wear filter gas masks. Take care to protect your skin and eyes. After inhalation, emergency response and treatment are similar to chlorine, hydrogen chloride and other irritant gases. If contacted with skin, first use paper or cotton to absorb the liquid, then rinse with water for at least 15 minutes; If Rinse with water, it can produce phosphoric acid and continue to burn the skin. Treat injury as acid burning. DANGER REGULATION GB8.1 Type 81040. Iron adjustment: level A corrosive inorganic acid product, 91022. a n. 1810. IMDG code Page 8197, TYPE 8. Oxychloride phosphorus is often used in industry industry There are many details to take note of when using oxychloride phosphorus to prevent security and accidents problems: 1. Preparation of oxychloride phosphorus: since many factories use water jet vacuum cleaner to store oxychloride phosphorus, great attention must be paid to the switch order to avoid accidents. The buffer vacuum tank must often be inspected to prevent water from being sucked into storage containers and causing accidents. 2. Oxychloride phosphorus with hydrogen chloride: take care not to add an excess of phosphorus chloride. If there is an excess, slowly fall into water to hydrolyze phosphorus chloride, then remove phosphorous acid produced. 3. Slowly place all materials containing phosphorus oxide in a hydrolysis kettle containing cold water and stir continuously. Avoid adding water to the material. 4. Whenever you use phosphorus-containing systems, take care to measure the water content in all components to prevent excess water from causing an explosion. 5. Frequently inspect crack and imperfections reactors to prevent leakage water and get in touch with Oxychloride phosphorus. 6. When adding phosphorus oxide to a reactor, they remain observant and stop the material immediately if there are abnormal events. It only keeps adding material after investigating the cause of these occurrences. 7. Oxychloride phosphorus reacts slowly with water when at lower temperature. So, slowly fall to prevent sudden intense reactions. 8. Oxychloride phosphorus can be used in combination with dichloroethan and toluene solutions, but not with alcohol and ammonia solutions. Category Articles Corrosive Toxicity Toxicity Toxicity Highly toxic Acute toxicity ACQUE-RAT LD50: 380 mg/kg Flammability and hazardous characteristics Being explosive Inventive through water with the release of toxic chloride, phosphorous gas storage and ventilated transport characteristics, at low temperature and dry; and store it separately from dry sand of alkali fire agent, dry stone powder; prohibit the use of TLV-TLV-TWA 0.1 PPM water standard (0.6 mg / m3); Stel 0.5 ppm (3 mg/m3) Oxychloride phosphorus is a clear liquid, colorless in yellow, smoking, fat with a pungent and moldy smell. Colorful smoke liquid with a pungent smell; density 1.645 g / ml; it blocks at 1 ° C; bubbles at 105.5 ° C; It reacts with water and ethanol. In the manufacture of pesticides, pharmaceuticals, plasticizers, petrol additives and hydraulic fluid. Oxychloride phosphorus is used to produce hydraulic fluids, plasticizers and fire agents; as a cloning agent; and as a solvent in criticism. Oxychloride phosphorus is an intermediateIn the production of foreign triarylphosphate (for example, trifenyl phosphate and trichesyl phosphate), which were used as flame retardants and PVC plasticizers. It is acutely toxic to the eyes, the throat and the respiratory tract. Phosphorus oxyloride is also used in in Reitrating, as a chlorination agent, in particular to replace oxygen in organic compounds, as solvent in the criticism and semiconductor industry. A white crystalline solid. It is a monobasic acid that forms the H2PO2 anion Ã ¢ â,¬ "in water. Sodium salt, and therefore acid, can be prepared by heating yellow phosphorus with sodium hydroxide solution. Free acid and its own Salts are powerful reducing agents. The oxychloride phosphorus can be prepared from the phosphorus trichloride or phosphorus pentachloride. It can be obtained from a phosphorus trichloride by adding potassium chlorate added: 3pcl3 + kclo3 Ã ¢ â € 3 POCL3 + KCL3 Ã ¢ â € 3 POCL3 + KCL Oxycloride is also obtained through the action of boric acid or oxalic acid with phosphorus pentacloride: 3pcl5 + 2B (ohÃ¯Â¼ â € â € â € œ € 3 POCL3 + B2O3 + 6HCL PCL5 + (CoohÃ¯Â¼ â € 2 â € œCl3 + Co + CO2 + 2HCL Phosphorus Oxychloride also made by calcium phosphate heating in a chlorine stream and carbon monoxide at 350 ° C: 2CA3 (PO4Ã¯Â¼ â € 2 + 9CL2 + 6CO Ã, â € 4 POCL3 + 6CACO3 alternatively, warming a mixture of calcium and carbon phosphate in A chlorine current at 750 Ã ° C produce oxycloride. A colorless smoke liquid with a pungent smell. DensitÃ 14.0 lb / gall. Very toxic inhalation and corrosive metal and fabric. Used in petrol additives and hydraulic fluids. The oxyloride phosphorus is reactive of water. Incompatible with strong oxidizing agents, alcohols, bases (including amines). It can vigorously or explosive react if mixed with Diisopropil ether or other ethers the P. resale of the quantities of traces of metal salts [J. Haz. Mat., 1981, 4, 291]. The combination of chloride with zinc powder has caused immediate ignition, due to the formation of phosphinal gas that lights up, [Mellor, 1940, vol. 8, 1025]. An exotherm that begins with the mixing of phosphorus oxychloride with acetone (a ketone) leads to an explosion, can behave in a similar way with other ketones, [research and development of the organic process, vol.4, n. 6.200, "Phosphorus Oxyloride and Acetone: an incompatibility investigation using Arc."] The mixture is very irritating to the skin, eyes and mucous membranes. Inhalation of his vapors can cause pulmonary edema. Inhalation of phosphorus vapors Ossicloride produced acute and chronic toxicity in test subjects. In humans, exposure to his vapors can cause headaches, dizziness, weakness, nausea, vomiting, cough, chest pain, bronchitis and pulmonary edema. Most of these symptoms manifest from chronic exposure to its vapors. LC50 Value, inhalation (rats): 48 ppm (301 mg / m3) / 4 hvapors of this compound are an irritant for eyes and mucous membranes. The corrosive liquid and can cause skin burns. An oral LD50 value for rats is documented to be 380 mg / kg (Niosh 1986) .. The poisonous, corrosive and irritating gases Generated when the oxychloride phosphorus is heated or in contact with water. The oxychloride phosphorus can turn on other combustible materials (wood, paper, oil, etc.). The oxychloride phosphorus reacts violently with water. When heated to decomposition, decomposition, Oxychloride emits toxic fumes of chlorides and phosphorus oxides; Phosphorus Oxychloride react with water or steam to produce toxic and corrosive fumes. Incompatible with carbon disulfide; N, n-dimethylformamide; 2.5-dimethylpyrrole; 2.6-dimethyl-oxide pyridine; dimethylsulfoxide; Ferrocene-1,1-acid-dicarboxylic; water; And zinc. Do not store combustible materials, in particular fibrous organic materials, or with electrical equipment or others that can be corroded. Reacts violently with moisture. Poison for inhalation and ingestion. A corrosive eye, the skin and the mucous membrane Irritant. The reaction potentially explosive with water evolves hydrogen chloride and phosphine, which then turns on. Explosive reaction with 2.6Dimethylpyridine N-oxide, dimethyl sulphoxide, ferrocene1, hydroxyl acid, n-oxide pyridne (over bo'c), sodmm + heat. Violent reaction or ignition with BI3, carbon disulfide, 2.5-Dimethyl pyrrete + formamide dimethil, organic matter, zinc powder. Reacts with water or steam to produce toxic and corrosive fumes. Incompatible with carbon disulfide, n, ndimethyl-formamide, 2.5-dunethylpyrrole, 2.6-dimethylpyridine n-oxide, dimethylsulfoxide, ferrocene1, I-dicarboxylic acid, water, zinc. When heated to decomposition emits highly toxic fumes of Cland Pox Phosphorus Oxychloride is used in the manufacture of pesticides, drugs, plasticizers, gasoline additives; and hydraulic fluids. UN1810 Phosphorus Oxychloride, Hazard class: 6.1; Labels: 6.1-plastic materials, 8-corrosive material, Hazard Zone B. distil The liquid under reduced pressure to separate it from most of the HCL and phosphoric acid (from hydrolysis); The intermediate fraction is redistelled to vials containing a little purified mercury. These vials are sealed and stored in the dark for 4-6 weeks with occasional agitation to facilitate the reaction of any free chloride with mercury. The POCL3 is then distiled fractionally and stored in vials sealed in the dark until it is required [Herber J am Chem Soc 82 792 1960]. Lewis and Sowerby [J Chem Soc 336 1957] Respussed their POCL3 distillate with na thread for 4 hours, then removed the na and distilled again. USA Na only with ICL3 almost pure to avoid explosions. Vapours Harmful; Work in a fume closet efficient. A powerful oxidant. Violently decomposes in water, forming warmth and acids hydrochlorides and phosphoric. Violent reaction with alcohols, phenols, amines, reducing agents; combustible materials; carbon disulfide; Dimethylformamide and many other materials. Rapid corrosion of metals, except nickel and lead. Pour on the sodium bicarbonate. Spray with aqueous ammonia and add crushed ice. Neutralize and pour into drain with running water. In accordance with 40CFR165, follow the recommendations for the disposal of pesticides and containers of It must be disposed of correctly following the indications of package labels or by contacting your local environmental or federal environmental control or contacting your regional EPA office. phosphorus

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